Copolymer electrode self-modified with fullerene C_{60}

Muhammed Mizher Radhi

Abstract
A copolymer of polystyrene-acrylonitrile was modified with fullerene C_{60}, using gamma irradiation technique to produce a new copolymer with high quality in physical and chemical properties and using to fabrication a working electrode alternative for the modification of commercial working electrode in cyclic voltammetric technique (such as glassy carbon electrode, Pt, Au electrode). The new copolymer electrode self-modified with C_{60} (CESMC_{60}) was characterized electrochemically by cyclic voltammetric method using a standard solution of 0.3 mM K_{4}[Fe(CN)_{6}] in 1 M KCl to evaluate the redox current peaks of Fe(II)/Fe(III). The electrochemical and physical properties of the new fabrication electrode has good hardness, insoluble in aqueous and non-aqueous solvent, electrochemical stability, resistance to high temperatures and different pH. It was found that 0.3 mM of K_{4}[Fe(CN)_{6}] in 1 M KCl of the new electrode at different concentration using reagent, scan rate, pH, temperature, electrolytes, stability and reliability. The new electrode was given encouraging properties for using in high-precision analysis and resuscitates commercial electrodes such as glassy carbon electrode, platinum electrode, gold electrode, and so on.

Keywords: cyclic voltammetry, copolymer electrode modified C_{60}, K_{4}[Fe(CN)_{6}], redox process

1. Introduction
Some scientists were manufactured high quality electrodes from conductive polymer which were used in the device of cyclic voltammetry and pH meter [1-3] and reference electrode [4].

IUPAC was upheld the modification of electrodes with conductive and semiconductive polymers to enhancement of the redox current peaks in cyclic voltammetric technique [5,6]. The modification of electrode in cyclic voltammetry was performed as chemical sensors and biosensors [7].

Fullerene C_{60} film as single wall on gold electrode was studied in cyclic voltammetric technique which characterization in this method [8,9]. Some polymer synthesis with nanoparticles like poly(methyl methacrylate) complex with Fullerenes C_{60} using as organic solar cell [10-12].

The use of C_{60} with the polymer composition has been gotten a polymer with different recipes for the polymer and that irradiation under vacuum pressure and at 300 °C to produce different spectral of conductive polymer [13]. C_{60} and carbon nanotubes (CNT) were deposed with the polymer to production of new recipes on the polymer where the polymer gave the new specification in photovoltaic cells manufacturing [14].

Some of researches were studied an application of important biosensors using fullerene C_{60}, the new biosensor electrode with nanomaterials which has a functionalization in electrochemical applications [15,16].

Polymer-fullerene was synthesized and characterized by cyclic voltammetric technique to observation of energetic offsets which has up to value of 150 meV, this energy contribute for improving performance of the polymer solar cells [17].

New electrochemical sensors were used in detection of ions in voltammetric analysis by modification of electrode with metalloporphyrin dimer – fullerene C_{60} [18].

Cyclic voltammetry was studied the modification glassy carbon electrode with single wall carbon nanotube and chitosan as a sensor to determination of acetaimophone, uric acid and ascorbic acid in human serum and urine medium with satisfactory results [19].

Cyclic voltammetry and differential pulse voltammetry were studied the modification gold electrode with fullerene-pyrrolidine to analysis the electrochemical characterized of the synthesized fullerene derivatives [20].

A thin film of C_{60} with nano-TiO_{2} on the electrode was determined the photocurrent efficiency photoactive of the polycrystalline TiO_{2} and decreases the recombination effect [21].

In this work, a new fabrication of copolymer electrode self-modified with C_{60} as alternative working electrode using in cyclic voltammetric technique was characterized.

2. Experimental
2.1. Synthesis of copolymer modified with fullerene C_{60}

The copolymer was manufactured process of polystyrene – acrylonitrile modified with Fullerene C_{60} using gamma-irradiation by ferrous ammonium sulfite (FAS) as a catalyst. The minimization of the copolymer production was the following:

- 0.25% (w/w) of the catalyst FAS
- 80% (w/w) of monomer (acrylonitrile)
- 1.5 MRD dose of gamma-irradiation
- 2 g of polymer (polystyrene)
- 1 mg of Fullerene C_{60}

Fig. 1 shows the mechanism reaction as a suggestion [22].
2.2. Fabrication of copolymer electrode self-modified with fullerene C$_{60}$ (CESMC$_{60}$)

The new fabrication working electrode (CESMC$_{60}$) has been fabricated by a diameter of 5 mm of the material of copolymer manufacturer, and entered into a hole works in the center of the piece of the copolymer restaurant and prove to the wire of platinum 1 cm length where it executes a very small part of the other party, and the other part links with copper wire and all parts covered with glass tube and prove with epoxy adhesive, as shown in Fig. 2.

2.3 Instrument and electrochemical analysis methods

EZstat series (potentiostat/galvanostat) NuVant Systems Inc. pioneering electrochemical technologies USA were used in this study. An Ag/AgCl (3 M NaCl) and platinum wire (1 mm diameter) were used as a reference and counter electrodes, respectively. The new fabrication working electrode CESMC$_{60}$ was used in this study.

Cyclic voltammetric cell was used in this technique by adding 10 ml of electrolyte in the quartz cell and immerse three electrodes in the supporting electrolyte, then all the electrodes was connected with potentiostat to finding the results of analysis by the cyclic voltammogram using personal computer as a soft ware of the analysis results.

2.4 Reagents

The following reagents were used: KCl powder purity from SCRC (China), potassium hexacyanoferrate II K$_4$[Fe(CN)$_6$]·3H$_2$O from Fluka AG Chem. (Fabrik CH-9470 Buchs Germany) and other chemicals in high purity from the manufactures. All reagents were prepared by deionized distilled water. All experimental analysis in cyclic voltammetry cell was dissolved oxygen remove with pure nitrogen gas bobbling for 10 to 15 min to avoid the peroxide producing from the oxidation of oxygen in the water of the solution in the cyclic voltammetric cell.

2.5. Scanning Electron Microscopy

Scanning Electron Microscopy (SEM) was used to examine the morphology of the copolymer of polystyrene-acrylonitrile and the copolymer modification with fullerene C$_{60}$ as shown in Fig. 3a and 3b, respectively.
3. Results and discussion

3.1. Electrochemical properties of the copolymer electrode self-modified with fullerene C$_{60}$

3.1.1. Effective of potential area

It can be compared to the potential area that are in commercial working electrode such as glassy carbon electrode (GCE), Platinum electrode (Pt-electrode), gold electrode (Au-electrode) with a new fabrication copolymer working electrode in 0.1 KCl electrolyte. It was noted that the new average potential area of the CESMC$_{60}$ is larger than GCE where the average potential area of CESMC$_{60}$ covers -1.8 to +2.0 V while the average potential area of GCE covers -1.5 to +1.8 V. It can also note that the presence of redox current peaks at commercial electrodes compared with flat area at CESMC$_{60}$.

3.1.2. Effect of CESMC$_{60}$ on the current and potential for redox current peaks in K$_4[Fe(CN)_6]$

Commonly it was used a compound of 1mM K$_4[Fe(CN)_6]$ in the calibration of the cyclic voltammetric technique, when using in aqueous solutions. Good redox current peaks Fe(III)/Fe(II) were found to the CESMC$_{60}$. Where it was noted that the separation potential between redox peaks value $\Delta E_{pc}$ produces zero current potential ($E_0, 1$) of 63mV for the process Fe(II)/Fe(III) at the new fabrication working electrode, the oxidation current peak of 63mV at low scan rate (10mV/sec) as described in the equation: $y=0.483x+2.1379$, $R^2=0.9976$.

From Fig. 6 the intercepts at zero current or scan rate produces zero current potential ($E_0, 1$) of 63mV for the process of Fe(II)/Fe(III) at CESMC$_{60}$ [25].

![Fig. 4 cyclic voltammogram of 0.3mM K$_4[Fe(CN)_6]$ in 1M KCl as a supporting electrolyte using CESMC$_{60}$ as working electrode versus Ag/AgCl as reference electrode at scan rate 100 mVs$^{-1}$](image)

**Table 1.** Scan rate, potential, current, peak potential separation, peak current ratio and diffusion coefficient ($D_f$) for Fe(II)/Fe(III) in KCl solution using CESMC$_{60}$

<table>
<thead>
<tr>
<th>Scan rate (Vsec$^{-1}$)</th>
<th>$E_{pc}$ (V)</th>
<th>$E_{pa}$ (V)</th>
<th>$I_{pc}$ (A)</th>
<th>$I_{pa}$ (A)</th>
<th>$\Delta E_{pc-Epa}$ (V)</th>
<th>Ratio $I_{pa}/I_{pc}$</th>
<th>$D_f$ (Cm$^{-2}$sec$^{-1}$)</th>
</tr>
</thead>
<tbody>
<tr>
<td>0.01</td>
<td>-0.274</td>
<td>0.052</td>
<td>0.194</td>
<td>0.144</td>
<td>0.222</td>
<td>0.74</td>
<td>1.59x10$^{-10}$</td>
</tr>
<tr>
<td>0.02</td>
<td>-0.338</td>
<td>0.111</td>
<td>0.0269</td>
<td>0.212</td>
<td>0.227</td>
<td>0.79</td>
<td>1.72x10$^{-10}$</td>
</tr>
<tr>
<td>0.03</td>
<td>-0.384</td>
<td>0.141</td>
<td>0.0320</td>
<td>0.259</td>
<td>0.243</td>
<td>0.81</td>
<td>1.71x10$^{-10}$</td>
</tr>
<tr>
<td>0.04</td>
<td>-0.419</td>
<td>0.170</td>
<td>0.0361</td>
<td>0.298</td>
<td>0.249</td>
<td>0.83</td>
<td>1.70x10$^{-10}$</td>
</tr>
<tr>
<td>0.05</td>
<td>-0.447</td>
<td>0.198</td>
<td>0.0394</td>
<td>0.327</td>
<td>0.249</td>
<td>0.83</td>
<td>1.64x10$^{-10}$</td>
</tr>
<tr>
<td>0.06</td>
<td>-0.471</td>
<td>0.217</td>
<td>-0.0424</td>
<td>0.356</td>
<td>0.254</td>
<td>0.84</td>
<td>1.62x10$^{-10}$</td>
</tr>
<tr>
<td>0.07</td>
<td>-0.498</td>
<td>0.238</td>
<td>-0.0447</td>
<td>0.381</td>
<td>0.260</td>
<td>0.85</td>
<td>1.59x10$^{-10}$</td>
</tr>
<tr>
<td>0.08</td>
<td>-0.509</td>
<td>0.247</td>
<td>-0.0472</td>
<td>0.406</td>
<td>0.262</td>
<td>0.86</td>
<td>1.58x10$^{-10}$</td>
</tr>
<tr>
<td>0.09</td>
<td>-0.534</td>
<td>0.263</td>
<td>-0.0493</td>
<td>0.427</td>
<td>0.271</td>
<td>0.87</td>
<td>1.55x10$^{-10}$</td>
</tr>
<tr>
<td>0.10</td>
<td>-0.542</td>
<td>0.279</td>
<td>-0.0515</td>
<td>0.446</td>
<td>0.263</td>
<td>0.87</td>
<td>1.52x10$^{-10}$</td>
</tr>
</tbody>
</table>

3.2. Effect of scan rate on the CESMC$_{60}$

The different scan rates (0.01 – 0.1 Vs$^{-1}$) was studied for the CESMC$_{60}$ in a solution of 1M KCl as a supported electrolyte and 0.3 mM K$_4[Fe(CN)_6]$. It has been observed that the redox current peaks of Fe(II)/Fe(III) increased with increasing the scan rate which get heterogeneous kinetics phenomenon as shown in Table 1 [23].

A rational linear relationship between Log($I_{pa}$) and Log(SR) of the CESMC$_{60}$ 0.3 mM K$_4[Fe(CN)_6]$ which described by the equation: $y=0.4835x-2.1379$, $R^2=0.9976$.

The slope of graph in Fig. 5 is 0.48 which is destination compatible with theoretical slope of 0.5 for diffusion controlled process [24].
3.3. Calculation of diffusion coefficient (Df)

The diffusion coefficient of the oxidation reaction for ferric ions to ferrous ions through cyclic voltammetry at the CESMC60 can be calculated by Eq. (1) and from the data in Table 1 [26,27]:

\[ I_{pa} = 2.69 \times 10^5 (n)^{3/2} A C_o (D_f)^{1/2} \nu^{1/2} \]  

(1)

Where:

- \( I_{pa} \) is the oxidative current peak,
- \( A \) is the area of the electrode,
- \( F \) is Faraday,
- \( C_o \) is the concentration,
- \( \nu \) is the scan rate and
- \( D_f \) is the diffusion coefficient of the reacting species.

It was found that the \( D_f \) values of Fe(II)/Fe(III) in the 0.1M KCl as supporting electrolyte has average value of \( D_f = 1.622 \times 10^{-10} \) cm²sec⁻¹ of the oxidative current peak of \( K_4[Fe(CN)_6] \) when used the new fabrication working electrode CESMC60, that means the heterogeneous transfer of electron through the electrolyte by nanomaterials fullerene C₆₀in the modified working electrode which acts as electro-catalyst for the enhancement of the processing the Fe(II)/Fe(III) in the electrolyte by increasing of scan rate against to the diffusion coefficient values as shown in Table 1 [28,29].

3.4. The influence of different temperatures on the CESMC₆₀

The influence of different temperatures on the new fabrication electrode was used the redox current peaks value of output potential using \( K_4[Fe(CN)_6] \) and increasing the temperature from 10 °C to 80 °C. The logarithm of current versus inverse temperature depends on the Arrhenius equations [24]:

\[ \sigma = \sigma_0 \exp \left( - \frac{E^*}{RT} \right) \]  

(2)

\[ D = D_0 \exp \left( - \frac{E^*}{RT} \right) \]  

(3)

Where:

- \( \sigma / D \) - is the quotient of connectivity
- \( \sigma_0 / D_0 \) -is the quotient of the standard connectivity
- \( E^* \) - is the activation energy
- \( R \) - is the ideal gas constant
- \( T \) - is the temperature in Kelvin

The activation energy which needed to oxidizing the Fe(II) to Fe(III) was found as \( E^* = 1.05 \) kJmol⁻¹K⁻¹ with less value than using other electrodes such as grafted polymer electrode in the same operation as \( E^* = 1.75 \) kJmol⁻¹K⁻¹ [30]. Also, the stability of GPE at high temperature has high resistance than other working electrodes [31-38].

3.5. Sensitivity of CESMC₆₀ in low detection limit at different concentrations

The new fabrication working electrode CESMC₆₀ of the polarization of high sensitivity in the detection of low concentrations of solutions of \( K_4[Fe(CN)_6] \) in 0.1M KCl solution as a supporting electrolyte, where it can be seen as a good sensor for low concentrations of chemical compounds with high sensitivity (R²=0.9822 for anodic current peak and R²=0.9834 for cathodic current peak) of the CESMC₆₀ for both oxidation-reduction current peaks as shown in Figs. 7 and 8 respectively.

3.6. Stability and reliability of CESMC₆₀

The potential cycling of the redox current peaks of the modified working electrode CESMC₆₀ in 0.3 mM \( K_4[Fe(CN)_6] \) with 1M KCl was carried out during cyclic voltammetry. Continuous potential cycling did not seem to affect the redox current peak of the new working electrode, since the faradic activity appears reliable even after ten times, and the relative standard deviation (RSD) of the cathodic current peak of the new fabrication electrode is ±4.16 % as shown in Table 2. Table 3 shows the reliability of CESMC₆₀ as working electrode at SR is 100 mV s⁻¹ for anodic current peak of 0.3 mM \( K_4[Fe(CN)_6] \) in 0.3 M KCl and the RSD is ±3.34 %.

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**Fig. 7** Plot of anodic current peak of different concentrations (0.3-1mM) \( K_4[Fe(CN)_6] \) in 1M KCl at fabrication working electrode CESMC₆₀ versus Ag/AgCl and SR 100 mVs⁻¹

**Fig. 8** Plot of cathodic current peak of different concentrations (0.3-1mM) \( K_4[Fe(CN)_6] \) in 1M KCl at fabrication working electrode CESMC₆₀ versus Ag/AgCl and SR 100 mVs⁻¹

**Fig. 9** Cyclic voltammogram of ten times for redox current peaks of CESMC₆₀ in 0.3 mM \( K_4[Fe(CN)_6] \) with 1M KCl at SR 100mVs⁻¹
Fig. 9 shows the cyclic voltammogram of redox current peaks of 0.3 mM K₄[Fe(CN)₆] in 1M KCl for ten times of Fe(II)/Fe(III) which illustrate the good stability of the cyclic voltammogram of the new fabrication working electrode by overlapping the lines.

<table>
<thead>
<tr>
<th>Numbers</th>
<th>Ipc (μA)</th>
<th>Mean</th>
<th>RSD</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>-49.0</td>
<td></td>
<td></td>
</tr>
<tr>
<td>2</td>
<td>-48.9</td>
<td></td>
<td></td>
</tr>
<tr>
<td>3</td>
<td>-47.9</td>
<td></td>
<td></td>
</tr>
<tr>
<td>4</td>
<td>-47.5</td>
<td></td>
<td></td>
</tr>
<tr>
<td>5</td>
<td>-46.9</td>
<td></td>
<td></td>
</tr>
<tr>
<td>6</td>
<td>-45.5</td>
<td></td>
<td></td>
</tr>
<tr>
<td>7</td>
<td>-45.8</td>
<td></td>
<td></td>
</tr>
<tr>
<td>8</td>
<td>-44.8</td>
<td></td>
<td></td>
</tr>
<tr>
<td>9</td>
<td>-44.4</td>
<td></td>
<td></td>
</tr>
<tr>
<td>10</td>
<td>-43.4</td>
<td>-46.41</td>
<td>±4.16%</td>
</tr>
</tbody>
</table>

Table 2. Reliability of CESMC₆₀ as working electrode at SR is 100 mVs⁻¹ for cathodic current peak of 0.3 mM K₄[Fe(CN)₆] in 1M KCl

Numbers Ipa (μA) Mean RSD
1 40.9
2 41.0
3 42.2
4 42.5
5 42.7
6 43.4
7 43.7
8 44.5
9 44.7
10 44.8 43.02 ±3.34%

Table 3. Reliability of CESMC₆₀ as working electrode at SR is 100 mVs⁻¹ for anodic current peak 0.3 mM K₄[Fe(CN)₆] in 1M KCl

3.6. Study the CESMC₆₀ at different pH

The new fabrication CESMC₆₀ was studied at different pH in aqueous solution for both acidic and alkaline medium as shown in Table 4. It was observed that the fabrication electrode properties of the redox current peaks for Fe(II)/Fe(III) in 0.3 mM K₄[Fe(CN)₆] with 1M KCl as a supporting electrolyte have the same effect against to the acidic and alkaline medium by increasing the current peak against to increasing of acidity and the same phenomena increasing the current peak versus ion increasing the alkaline medium as shown in Table 4 and Figs. 10 and 11.

<table>
<thead>
<tr>
<th>pH</th>
<th>Epc V</th>
<th>Epa V</th>
<th>Ipc mA</th>
<th>Ipa mA</th>
</tr>
</thead>
<tbody>
<tr>
<td>3</td>
<td>-544</td>
<td>263</td>
<td>-47.2</td>
<td>41.3</td>
</tr>
<tr>
<td>4</td>
<td>-531</td>
<td>228</td>
<td>-60.4</td>
<td>41.9</td>
</tr>
<tr>
<td>5</td>
<td>-583</td>
<td>238</td>
<td>-49</td>
<td>41.3</td>
</tr>
<tr>
<td>6</td>
<td>-574</td>
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<td>-51.2</td>
<td>41.9</td>
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<tr>
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<td>40</td>
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<tr>
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<td>347</td>
<td>-55.4</td>
<td>48.1</td>
</tr>
<tr>
<td>11</td>
<td>-580</td>
<td>317</td>
<td>-48.6</td>
<td>42.3</td>
</tr>
</tbody>
</table>

Table 4. Relationship between pH and redox current and potential peaks of 0.3 mM K₄[Fe(CN)₆] with 1M KCl at CESMC₆₀

4. Conclusions

A new fabrication working electrode CESMC₆₀ was manufactured and characterized electrochemically in the laboratory of cyclic voltammetry. All chemical and physical properties was studied using 0.3 mM K₄[Fe(CN)₆] in 1M KCl for redox current peaks of reversible reaction of Fe(II)/Fe(III). It was found an enhancement current peak of the CESMC₆₀ for redox process of Fe(II)/Fe(III) was observed compared to the GCE about ten times, also extended potential area working region between 2.0 to -2.0 V for the new electrode. The sensitivity of the fabrication electrode has high detection analysis using cyclic voltammetry significantly depend on the different concentration, pH, temperature, scan rates, and Stability and reliability. It was determined the diffusion coefficient (Dₜ) of the CESMC₆₀ in the 0.3 mM K₄[Fe(CN)₆] in 1M KCl which has average value 1.6x10⁻¹⁰ cm²sec⁻¹.

References


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**Ref:**